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Problem **Why is the Sky Blue?** Find the

Average Atomic Mass
- Example:

Magnesium

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SECTION 12.1 THE

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ARITHMETIC OF EQUATIONS (pages 353–358) This section explains how to calculate the amount of reactants required or product formed in a nonchemical process. It teaches you how to interpret chemical equations in terms of interacting moles, representative particles, masses,

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and gas volume at
STP.

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(Stoichiometry) p. 357

#2 p. 379 #61, 64, 69,
70, 73, 86, 88, 90 p.

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5-10 p. 880 Chapter

14 #22 Answers:

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SECTION 12.1 THE ARITHMETIC OF EQUATIONS (pages 353–358) This section explains how to calculate the amount of reactants required or product formed in a nonchemical process.

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SECTION 12.1 THE

ARITHMETIC OF

EQUATIONS (pages

353–358) This section

explains how to

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of reactants required

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or product formed in a nonchemical process. It teaches you how to interpret chemical equations in terms of interacting moles, representative particles, masses, and gas volume at STP.

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SECTION 12.1 THE
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EQUATIONS (pages
353–358) This section
explains how to
calculate the amount
of reactants required
or product formed in a
nonchemical process.

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Reading Answers

Introduce the term sto- ichiometry in your own words. Stress that stoichiometry allows students to calculate the amounts of chemical sub- stances involved in chemical reactions using information

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obtained from balanced chemical equations.

12.1 The Arithmetic of Equations 12

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Section 3.1 – 3.2 1.

True or False? Most hydrogen atoms have a mass of 1.008 amu. Justify your answer. If true, explain why it is true. If false, what mass do most

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amu is actually
hydrogen's average
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