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practice, the amount of product obtained is called the actual yield, and it is often less than the theoretical yield for a number of reasons.

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theoretical yield, expressed as a percentage.

$$\text{Percent Yield} = \frac{\text{Actual}}{\text{Theoretical}} \times 100\%$$

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Answers percent yield of the reaction when 77.0 g of CO_2 are formed from burning 2.00 moles of C_5H_{12} in 4.00 moles of O_2 .

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The theoretical yield is what you calculate when you do a calculation on paper or before you do a reaction in a lab. The actual yield will always be less than the theoretical yield because no chemical reaction ever reaches 100 percent completion. In a lab setting, there ' s always

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some amount of error, whether it ' s big or small.

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Now, we use the actual
yield and the theoretical
yield to calculate the
percent yield. Step 1: List
the known quantities and
plan the problem.

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Known. Actual yield =
14.9 g; Theoretical yield
= 15.7 g (from Part
12.11A) Unknown.
Percent yield = ?

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diagram and the example problem. Example Problem: The following chemical equation represents the production of gallium oxide, a substance used in the manufacturing of some semiconductor devices.

$$4\text{Ga}(s) + 3\text{O}_2(g) \rightarrow 2\text{Ga}_2\text{O}_3(s)$$

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Technology

Your company currently uses a process with a similar cost of materials that has an average percent yield of 91 percent. If the average percent yield of this process is higher than that, this could save the company money.

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