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reaction CHEMISTRY - Dynamic
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Equilibrium: Crash Course
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equilibrium by Thaddeus Mbaluka

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7.1 Dynamic equilibrium (SL)

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2 Chemistry

Reversible reactions and dynamic equilibrium

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Answers: Key The following equation describes a system that is at equilibrium: $2\text{H}_2(\text{g}) + \text{O}_2(\text{g}) \rightleftharpoons 2\text{H}_2\text{O}(\text{g}) + \text{energy (heat)}$ In Table 1 apply Le Chatelier's Principle and indicate the direction of the shift in equilibrium if the

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Quarantine Changed Our Whole
Relationship Dynamic These gases
react as follows: $A(g) + 3 B(g)$

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$C(g) + 2 D(g)$ At equilibrium, the container has 0.4 moles of D. a) Calculate the moles of A, B, and C in

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Answer Key At equilibrium the concentrations of reactants and products is still changing, however, the rate of the forward reaction (k_f) is equal to the rate of the reverse reaction (k_r) in what is described as a dynamic equilibrium such that. Page 6/9.

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Reversible reactions never come to an end: they just reach equilibrium.

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Answers at equilibrium: $2\text{H}_2 (\text{g}) + \text{O}_2 (\text{g}) \rightleftharpoons 2\text{H}_2\text{O} (\text{g}) + \text{energy (heat)}$

In Table 1 apply Le Chatelier's Principle and indicate the direction of the shift in equilibrium if the

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indicated stress is applied to the reaction system. Mrs. Zuberbuehler - Mrs. Zuberbuehler

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equilibrium if the indicated stress is applied to the reaction system. (The first one is completed for you.) Key Questions 1. Complete the following table:

Mrs. Zuberbuehler - Mrs.

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Zuberbuehler

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